

THE ULTIMATE GENERAL CHEMISTRY 1 STUDY GUIDE

- ❖ Matter and Measurement
 - Matter: anything that has mass and occupies space
 - Properties of matter: physical properties and chemical properties
 - States of matter: solid, liquid, gas
 - Measurements: SI Units, significant figures, dimensional analysis
- ❖ Atoms, Molecules, and Ions
 - Atomic structure: protons, neutrons, and electrons
 - Periodic table and trends: electronegativity, ionization energy, electron affinity
 - Chemical bonding: ionic bonding, covalent bonding, metallic bonding
 - Chemical formulas and equations: empirical and molecular formulas, balancing chemical equations
- ❖ Stoichiometry
 - Limiting reactants and percent yield
 - Mole concept and Avogadro's number
 - Molarity and molality
 - Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
- ❖ Chemical Reactions and Equations
 - Types of chemical reactions: synthesis, decomposition, single replacement, double replacement, combustion
 - Redox reactions and balancing redox equations
 - Reaction rates: factors affecting reaction rates, rate laws, and reaction mechanisms
- ❖ Gases
 - Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
 - Kinetic Molecular Theory of gases: pressure, temperature, volume, and the number of moles of gas particles
 - Natural gases: van der Waals equation

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Martin Silberberg

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