

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



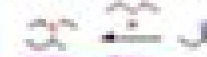
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding



Bonding energy



There are two types of bonding

Definition of ionic bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two double bonds** consist of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of atoms, covalent bonds can be polar and nonpolar.

• **Polar covalent bond** is formed when the difference in electronegativity is between 0.5 and 1.7.

• **Nonpolar covalent bond** is formed when the difference in electronegativity is less than 0.5.

• **Ionic bond** is formed when the difference in electronegativity is greater than 1.7.

• **Coordinate covalent bond** is formed when one atom provides both electrons for the bond.

• **Hydrogen bond** is a weak intermolecular force.

• **Van der Waals force** is a weak intermolecular force.

• **London dispersion force** is a weak intermolecular force.

• **Dipole-dipole interaction** is a weak intermolecular force.

• **Ion-dipole interaction** is a weak intermolecular force.

• **Ion-ion interaction** is a weak intermolecular force.

• **Hydrophobic interaction** is a weak intermolecular force.

• **Hydrophilic interaction** is a weak intermolecular force.

• **Hydrophobic effect** is a weak intermolecular force.

• **Hydrophilic effect** is a weak intermolecular force.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The direction of the dipole moment is from the less electronegative atom to the more electronegative atom.

• The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

• **δ⁻ (delta minus)** indicates an electron-deficient atom.

• **δ⁺ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. If the sum is zero, the molecule is nonpolar. If the sum is not zero, the molecule is polar.



Water is a polar molecule because the dipole moments of the O-H bonds do not cancel out.



Carbon dioxide is a nonpolar molecule because the dipole moments of the C=O bonds cancel out.

Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds because they have more electron density between the nuclei.

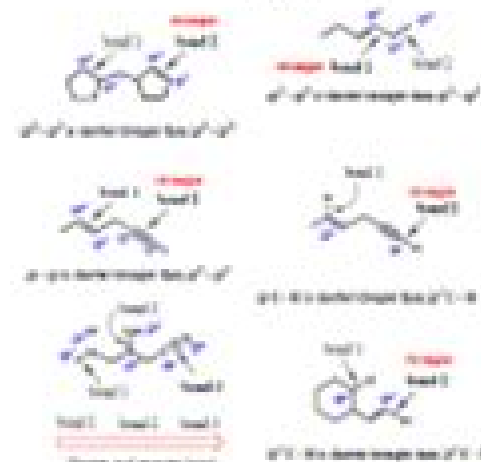


The length of a bond is the distance between the nuclei of the two atoms.

	sp ³	sp ²	sp
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	121 pm
Strength	338 kJ/mol	472 kJ/mol	590 kJ/mol
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	121 pm
Strength	338 kJ/mol	472 kJ/mol	590 kJ/mol

Higher s character means stronger bond

Example



Study Guide For Organic Chemistry

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