

100

Identify the independent variable, the dependent variable, and the control group. (2008) (1 point) correctly identifies each

1. **Q** 10. If the energy content of the chemical components is expressed as positive or negative values?
   
 a. **A** ☐ Positive
   
 b. **A** ☒ Negative
   
 c. **A** ☐ Both
   
 d. **A** ☐ None
2. **Q** 10. For the chemical  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ , which factor represents enthalpy of combustion?
   
 a. **A** ☒  $\Delta H_{\text{CH}_4}$ 
  
 b. **A** ☐  $\Delta H_{\text{O}_2}$ 
  
 c. **A** ☐  $\Delta H_{\text{CO}_2}$ 
  
 d. **A** ☐  $\Delta H_{\text{H}_2\text{O}}$
3. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
4. **Q** 10. For the chemical  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ , which factor represents enthalpy of combustion?
   
 a. **A** ☐  $\Delta H_{\text{CH}_4}$ 
  
 b. **A** ☒  $\Delta H_{\text{O}_2}$ 
  
 c. **A** ☐  $\Delta H_{\text{CO}_2}$ 
  
 d. **A** ☐  $\Delta H_{\text{H}_2\text{O}}$
5. **Q** 10. Which is the state of the water in the following reaction?  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ 
  
 a. **A** ☐ Gaseous
   
 b. **A** ☒ Liquid
   
 c. **A** ☐ Solid
   
 d. **A** ☐ None
6. **Q** 10. Which is the state of the water in the following reaction?  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ 
  
 a. **A** ☐ Gaseous
   
 b. **A** ☒ Liquid
   
 c. **A** ☐ Solid
   
 d. **A** ☐ None
7. **Q** 10. A compound's standard enthalpy of formation is 100 kJ/mol. Which statement is the correct one?
   
 a. **A** ☐ 100 kJ/mol
   
 b. **A** ☒ -100 kJ/mol
   
 c. **A** ☐ 100 kJ/mol
   
 d. **A** ☐ -100 kJ/mol
8. **Q** 10. For the chemical  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ , which factor represents enthalpy of combustion?
   
 a. **A** ☒  $\Delta H_{\text{CH}_4}$ 
  
 b. **A** ☐  $\Delta H_{\text{O}_2}$ 
  
 c. **A** ☐  $\Delta H_{\text{CO}_2}$ 
  
 d. **A** ☐  $\Delta H_{\text{H}_2\text{O}}$
9. **Q** 10. For the chemical  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ , which factor represents enthalpy of combustion?
   
 a. **A** ☐  $\Delta H_{\text{CH}_4}$ 
  
 b. **A** ☒  $\Delta H_{\text{O}_2}$ 
  
 c. **A** ☐  $\Delta H_{\text{CO}_2}$ 
  
 d. **A** ☐  $\Delta H_{\text{H}_2\text{O}}$
10. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
11. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
12. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
13. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
14. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
15. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
16. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
17. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
18. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
19. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
20. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
21. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
22. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
23. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
24. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
25. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
26. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d. **A** ☐ Heat of the substance
27. **Q** 10. The enthalpy of combustion is the amount of
   
 a. **A** ☐ Heat of the substance
   
 b. **A** ☒ Heat released during
   
 c. **A** ☐ Heat of the substance
   
 d.

# Multiple Choice Stoichiometry Test

**EW Minium**



## Multiple Choice Stoichiometry Test:

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