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MATHEMATICS P1

May/June 2023

MARKS: 150  
TIME: 3 hours

The question paper consists of 8 pages and 1 reference sheet.



$\tan \theta = \frac{\sin \theta}{\cos \theta}$

$E = mc^2$

$a^2 + b^2 = c^2$

$2(1 \times 8)$

$1 - 2 \approx 0$

$g = 9.81 \text{ m/s}^2$

$A$

$P(A \cap B)$

$3! = 3 \times 2 \times 1$

$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$

$\infty$

**FULL MEMO**

**GRADE 12**

**MATHEMATICS**

**PAPER 1**

**NSC JUNE 2023**

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# Question Paper 1 Topics Chapter For Examination

## Mathematics Grade 12

**C.F. HODGSON**

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3 655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1 1 we could experimentally determine the

number of moles of the acid in the solution to be 000731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $k_a$  or  $k_b$  from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $k_{a1}$  and  $k_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1 hcl aq naoh aq nacl aq h<sub>2</sub>o 1 hcl a q naoh a q nacl a q h 2 o l one mole of hcl hcl would be fully neutralized by one mole of naoh naoh

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance h<sub>2</sub>c<sub>2</sub>o aq 2naoh aq na<sub>2</sub>c<sub>2</sub>o<sub>4</sub> 4h<sub>2</sub>o to differentiate an acid from a base in a solution a ph scale is

used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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