



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

### Key Vocabulary

|   |   |
|---|---|
| <ul style="list-style-type: none"><li>hydrogen ion</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li><li>pH</li></ul> | <ul style="list-style-type: none"><li>hydrogen ion concentration</li><li>acidic</li><li>basic</li><li>neutral</li><li>titration</li><li>titration curve</li></ul> |
|---|---|

### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Diagrams

- Titration curve
- Titration curve

### Points of Consideration

Acidic and basic properties are determined by the concentration of hydrogen ions in a solution. The concentration of hydrogen ions is measured by the pH of a solution.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of other ions in a solution is determined by the concentration of the ions in the solution.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of other ions in a solution is determined by the concentration of the ions in the solution.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of other ions in a solution is determined by the concentration of the ions in the solution.

### Key Equations

|                                      |                                      |
|--------------------------------------|--------------------------------------|
| $\text{pH} = -\log_{10}[\text{H}^+]$ | $\text{pH} = \log_{10}[\text{OH}^-]$ |
| $\text{pH} = 7$                      | $\text{pH} = 14$                     |
| $\text{pH} = 1$                      | $\text{pH} = 13$                     |
| $\text{pH} = 2$                      | $\text{pH} = 12$                     |
| $\text{pH} = 3$                      | $\text{pH} = 11$                     |
| $\text{pH} = 4$                      | $\text{pH} = 10$                     |
| $\text{pH} = 5$                      | $\text{pH} = 9$                      |
| $\text{pH} = 6$                      | $\text{pH} = 8$                      |
| $\text{pH} = 7$                      | $\text{pH} = 7$                      |
| $\text{pH} = 8$                      | $\text{pH} = 6$                      |
| $\text{pH} = 9$                      | $\text{pH} = 5$                      |
| $\text{pH} = 10$                     | $\text{pH} = 4$                      |
| $\text{pH} = 11$                     | $\text{pH} = 3$                      |
| $\text{pH} = 12$                     | $\text{pH} = 2$                      |
| $\text{pH} = 13$                     | $\text{pH} = 1$                      |
| $\text{pH} = 14$                     | $\text{pH} = 0$                      |

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