

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$ | $pOH = -\log[OH^-]$ |

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, and give their K_a and K_b values.

1. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$

2. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$

3. Hydrochloric acid, HCl , $K_a = 1.0 \times 10^7$

4. Hydrobromic acid, HBr , $K_a = 1.0 \times 10^7$

5. Hydroiodic acid, HI , $K_a = 1.0 \times 10^7$

6. Nitric acid, HNO_3 , $K_a = 2.4 \times 10^{-4}$

7. Sulfuric acid, H_2SO_4 , $K_a = 1.0 \times 10^{-2}$

8. Phosphoric acid, H_3PO_4 , $K_a = 7.5 \times 10^{-3}$

9. Carbonic acid, H_2CO_3 , $K_a = 4.3 \times 10^{-7}$

10. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$

11. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$

1. Acetic acid, CH_3COOH
2. Hydrofluoric acid, HF
3. Hydrochloric acid, HCl
4. Hydrobromic acid, HBr
5. Hydroiodic acid, HI
6. Nitric acid, HNO_3
7. Sulfuric acid, H_2SO_4
8. Phosphoric acid, H_3PO_4
9. Carbonic acid, H_2CO_3
10. Acetic acid, CH_3COOH
11. Hydrofluoric acid, HF

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