

**Chemistry 2<sup>nd</sup> Semester Exam Review Sheet:****Chapters 9, 10, 11, 12, 13, 14, 15, and 16****Chapter 9: Chemical Quantities**

1. Solve stoichiometry problems using grams  $\rightarrow$  moles  $\rightarrow$  moles  $\rightarrow$  grams calculations. Examples below:  
a. Calculate moles/grams of  $(\text{NH}_4)_2\text{SO}_4$  produced when 30.0 moles of  $\text{NH}_3$  react with  $\text{H}_2\text{SO}_4$ .



$$30.0 \text{ moles} \times 1 \text{ mol/2 mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ g/mol/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b.  $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$ . How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams  $\rightarrow$  Moles  $\rightarrow$  Moles  $\rightarrow$  Grams

$$100. \text{ g } \text{Na}_2\text{CO}_3 \times 1 \text{ mol/106.0 g} \times 2 \text{ mol } \text{NaNO}_3 / 1 \text{ mol } \text{Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g } \text{NaNO}_3 \times 1 \text{ mol/85.00 g} \times 2 \text{ mol / 2 mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of  $\text{NaNO}_3$  if 100.0 grams of  $\text{Na}_2\text{CO}_3$  are used in the reaction? What is the percent yield of  $\text{NaNO}_3$  if 130.5 grams of  $\text{NaNO}_3$  are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

**Chapter 10: Modern Atomic Theory**

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? Absorb a photon of energy.

How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms (page 285-86)

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, proba 54, 96

# 2nd Semester Exam Chemistry Form

**Richard Bailey**

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**Grade Booster Question Bank Semester 2 Chemistry class 10** Amar Bhutani,2025-09-22 The Grade Booster Semester 2 Chemistry Question Bank for Class 10 equips students with all the essential practice they need for board examinations. It provides chapter based solved questions, important chemical equations, and well-illustrated diagrams to make complex concepts easier to understand. Examiner's guidelines and detailed answers ensure that students learn how to present responses effectively in the exam. The book also highlights important reactions, definitions, and formula-based questions that are frequently asked in exams. With focused revision support and ample practice material, this book helps students strengthen their conceptual clarity and enhances their ability to score high in ICSE Semester 2 Chemistry.

**Oswal - Gurukul Chemistry Most Likely Question Bank : ICSE Class 9 Sem II Exam 2022** Oswal - Gurukul,2022-01-09    **The College Blue Book** Huber William Hurt, Harriett-Jeanne Hurt, 1969    **Goyal's I.C.S.E Chemistry Question Bank with Model Test Papers For Class X Edition 2021** Goyal Brothers Prakashan,2021-09-01 CISCE's Modified Assessment Plan for Academic Year 2021-22 Reduced and Bifurcated Syllabus for First Semester Examination Chapterwise Important Points Chapter wise Multiple Choice Questions Specimen Question Paper issued by the CISCE 5 Model Test Papers based on the latest specimen question paper for First Semester Examination to be held in November 2021 Goyal Brothers Prakashan    **The Medical times and gazette** ,1881    **British Medical Journal** ,1890    Education Outlook ,1928    *"The" Lancet* ,1900    **Instructional Television Research** Pennsylvania State University. Instructional Film Research Program,1955    **The Medical circular [afterw.] The London medical press & circular [afterw.] The Medical press & circular** ,1872    Final Report: Sources and documentation United States. Advisory Committee on Human Radiation Experiments,1995    Program Bryn Mawr College,1898    **University of Iowa Studies in Education** ,1930    **Oxford University Gazette** University of Oxford,1894    **A Study of Interests and Their Relations to Other Factors of Achievement in the Elementary School Subjects** A. V. Smith, Edward Joseph Westenberger, Marie Cecelia McGrath, Mary Callixta, Paul Hanly Furley, Robert Thomas Rock, Sister Marie Cecilia Mangold, Sister Mary Columba, Sister Mary Immaculata, Sister Mary Irmina, Sylvester Schmitz, Thomas George Foran, William Dollard Commins, 1925    **Report on medical education ... 1889-91** ,1889    **Educational Research Monographs** ,1927    *The Content and Measurement of the Reading Vocabulary* Sister Mary Irmina,1927    Educational Research Bulletins ,1927    *Educational Research Monographs* Catholic University of America,1927

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