

Chemistry Study Guide

Periodic Table of the Elements

H 1 Hydrogen																	He 2 Helium
Li 3 Lithium	Be 4 Beryllium											B 5 Boron	C 6 Carbon	N 7 Nitrogen	O 8 Oxygen	F 9 Fluorine	Ne 10 Neon
Na 11 Sodium	Mg 12 Magnesium											Al 13 Aluminum	Si 14 Silicon	P 15 Phosphorus	S 16 Sulfur	Cl 17 Chlorine	Ar 18 Argon
K 19 Potassium	Ca 20 Calcium	Sc 21 Scandium	Ti 22 Titanium	V 23 Vanadium	Cr 24 Chromium	Mn 25 Manganese	Fe 26 Iron	Co 27 Cobalt	Ni 28 Nickel	Cu 29 Copper	Zn 30 Zinc	Ga 31 Gallium	Ge 32 Germanium	As 33 Arsenic	Se 34 Selenium	Br 35 Bromine	Kr 36 Krypton
Rb 37 Rubidium	Sr 38 Strontium	Y 39 Yttrium	Zr 40 Zirconium	Nb 41 Niobium	Mo 42 Molybdenum	Tc 43 Technetium	Ru 44 Ruthenium	Rh 45 Rhodium	Pd 46 Palladium	Ag 47 Silver	Cd 48 Cadmium	In 49 Indium	Sn 50 Tin	Sb 51 Antimony	Te 52 Tellurium	I 53 Iodine	Xe 54 Xenon
Cs 55 Cesium	Ba 56 Barium		Hf 72 Hafnium	Ta 73 Tantalum	W 74 Tungsten	Re 75 Rhenium	Os 76 Osmium	Ir 77 Iridium	Pt 78 Platinum	Au 79 Gold	Hg 80 Mercury	Tl 81 Thallium	Pb 82 Lead	Bi 83 Bismuth	Po 84 Polonium	At 85 Astatine	Rn 86 Radon
Fr 87 Francium	Ra 88 Radium		Rf 104 Rutherfordium	Db 105 Dubnium	Sg 106 Seaborgium	Bh 107 Bohrium	Hs 108 Hassium	Mt 109 Meitnerium	Ds 110 Darmstadtium	Rg 111 Roentgenium	Cn 112 Copernicium	Uut 113 Ununtrium	Fl 114 Flerovium	Uup 115 Ununpentium	Lv 116 Livermorium	Uus 117 Ununseptium	Uuo 118 Ununoctium
		<div> <div> Lanthanum 57 La </div> <div> Cerium 58 Ce </div> <div> Praseodymium 59 Pr </div> <div> Neodymium 60 Nd </div> <div> Promethium 61 Pm </div> <div> Samarium 62 Sm </div> <div> Europium 63 Eu </div> <div> Gadolinium 64 Gd </div> <div> Terbium 65 Tb </div> <div> Dysprosium 66 Dy </div> <div> Holmium 67 Ho </div> <div> Erbium 68 Er </div> <div> Thulium 69 Tm </div> <div> Ytterbium 70 Yb </div> <div> Lutetium 71 Lu </div> </div>															
		<div> <div> Actinium 89 Ac </div> <div> Thorium 90 Th </div> <div> Protactinium 91 Pa </div> <div> Uranium 92 U </div> <div> Neptunium 93 Np </div> <div> Plutonium 94 Pu </div> <div> Americium 95 Am </div> <div> Curium 96 Cm </div> <div> Berkelium 97 Bk </div> <div> Californium 98 Cf </div> <div> Einsteinium 99 Es </div> <div> Fermium 100 Fm </div> <div> Mendelevium 101 Md </div> <div> Nobelium 102 No </div> <div> Lawrencium 103 Lr </div> </div>															

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

H Kauffman



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Table of Contents Section 2nuclear Chemistry Study Guide Key

1. Understanding the eBook Section 2nuclear Chemistry Study Guide Key
 - The Rise of Digital Reading Section 2nuclear Chemistry Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 2nuclear Chemistry Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 2nuclear Chemistry Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2nuclear Chemistry Study Guide Key
 - Personalized Recommendations
 - Section 2nuclear Chemistry Study Guide Key User Reviews and Ratings

- Section 2nuclear Chemistry Study Guide Key and Bestseller Lists
- 5. Accessing Section 2nuclear Chemistry Study Guide Key Free and Paid eBooks
 - Section 2nuclear Chemistry Study Guide Key Public Domain eBooks
 - Section 2nuclear Chemistry Study Guide Key eBook Subscription Services
 - Section 2nuclear Chemistry Study Guide Key Budget-Friendly Options
- 6. Navigating Section 2nuclear Chemistry Study Guide Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 2nuclear Chemistry Study Guide Key Compatibility with Devices
 - Section 2nuclear Chemistry Study Guide Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 2nuclear Chemistry Study Guide Key
 - Highlighting and Note-Taking Section 2nuclear Chemistry Study Guide Key
 - Interactive Elements Section 2nuclear Chemistry Study Guide Key
- 8. Staying Engaged with Section 2nuclear Chemistry Study Guide Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 2nuclear Chemistry Study Guide Key
- 9. Balancing eBooks and Physical Books Section 2nuclear Chemistry Study Guide Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 2nuclear Chemistry Study Guide Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 2nuclear Chemistry Study Guide Key
 - Setting Reading Goals Section 2nuclear Chemistry Study Guide Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 2nuclear Chemistry Study Guide Key
 - Fact-Checking eBook Content of Section 2nuclear Chemistry Study Guide Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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