

Physical Chemistry Formula Revision

Basic Concepts

$$\text{Average atomic mass} = \frac{a_1 p_1 + a_2 p_2 + a_3 p_3 + \dots}{100}$$

 a_1, a_2, a_3, \dots = atomic mass of isotopes. x_1, x_2, x_3, \dots = percentage abundances

Male concept

$$\text{no. of males} = n = \frac{\text{given mass}}{\text{avg mass}} = \frac{91}{21}$$

$$n = \frac{\text{no. of particles}}{\text{Avogadro's no.}} = \frac{N}{6.023 \times 10^{23}}$$

$$n = \frac{\text{vol. of gas}}{\text{vol. of gas at STP}} = \frac{V}{22.4 \text{ l}}$$

Concentration Times

$$\text{Percentage, \%} = \frac{W_2 \text{ or } V_2}{W_1 \text{ or } V_1} \times 100$$

$$\text{Molarity, } M = \frac{n_b}{V(l)} = \frac{N_b \times 1000}{M_b \times V(ml)}$$

$$\text{Normality, } N = \frac{n_{\text{eq}}}{V(l)} = \frac{M_b \times 1000}{E_b \times V(\text{ml})}$$

$\hookrightarrow \frac{M_b}{n_f}$

parts per million, ppm = $\frac{W_b}{W} \times 10^6$

$$\text{Molality, } m = \frac{n_b}{M_b (\text{kg})} = \frac{M_b \times 1000}{M_b \times M_b (\text{g})}$$

Mole fraction, $x_b = \frac{n_b}{n_a + n_b + n_c}$

Stoichiometry



$$\frac{\text{moles of A}}{a} = \frac{\text{moles of B}}{b} = \frac{\text{moles of C}}{c}$$

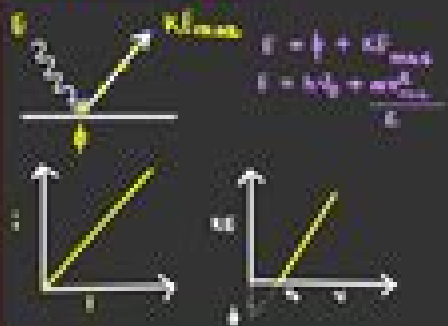
Limiting R_p : $\frac{\text{given amt.}}{\text{stoichiometric coeff.}} \rightarrow \text{least}$

Atomic Structure

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$$E = \frac{hc}{\lambda} \quad \text{and} \quad \lambda = \frac{12400}{E(\text{eV})}$$

Maths: Electrons & Energy



The Giving Circle: Why you should

$$m_e h = \frac{h}{\lambda} \rightarrow \lambda = \frac{h}{m_e} \quad \lambda = \frac{h}{\sqrt{2 m_e E}}$$

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$$\text{average} = \frac{m \cdot h_1}{E \cdot x} \quad \text{where } m = 0.5 \times \frac{m^2}{E} \cdot h^2$$

www.ck12.org

$$E_{\text{eff}} = E_0 \left(1 + \frac{E_0}{E_c} \right) \quad \text{for } E_0 \ll E_c$$

$$\frac{1}{\lambda} = \frac{1}{\lambda_0} \left[\frac{1}{\sqrt{1 - \frac{v^2}{c^2}}} \right] \quad \lambda_0 = 1000 \text{ nm}$$

max. no. of spectral lines when λ^{th} electron from
 n_2 to n_1 : $\frac{(n_2 - n_1)(n_2 - n_1 + 1)}{2}$

Hilbert's Uniformity Principle: $\forall \epsilon > 0 \exists \frac{h}{4\pi}$

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total nodes = $n-1$

[illegible]

$$\text{modified_nodes} = m - 1 - l$$

Keywords: *work engagement, organizational commitment, turnover intentions, organizational citizenship behaviors, organizational trust*

$$\text{control} = \sqrt{1 - \frac{1}{N}} \approx 0.99$$

Figure 1

Structure Of Physical Chemistry

Cyril Norman Hinshelwood



Structure Of Physical Chemistry:

The Structure of Physical Chemistry Cyril Norman Hinshelwood,1951 **The Structure of Physical Chemistry** C. N. Hinshelwood,1951 *The Structure of Physical Chemistry* Cyril Norman Hinshelwood,1951 **The Structure of Physical Chemistry** Sir Cyril Norman Hinshelwood,2005 The Structure Of Physical Chemistry Hinselwood C. N,2022-10-27 This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant *Principles of Physical Chemistry* Hans Kuhn,Horst-Dieter Försterling,David H. Waldeck,2009-03-17 *Principles of Physical Chemistry* Second Edition uniquely uses simple physical models as well as rigorous treatments for understanding molecular and supramolecular systems and processes In this way the presentation assists students in developing an intuitive understanding of the subjects as well as skill in quantitative manipulations The unifying nature of physical chemistry is emphasized in the book by its organization beginning with atoms and molecules and proceeding to molecular assemblies of increasing complexity ending with the emergence of matter that carries information i e the origin of life a physicochemical process of unique importance The aim is to show the broad scope and coherence of physical chemistry **An Introduction to the Principles of Physical Chemistry from the Standpoint of Modern Atomistics and Thermo-dynamics** Edward Wight Washburn,1915 *A System of Physical Chemistry* William Cudmore McCullagh Lewis,1916 *Physical Chemistry and the Structure of Matter* David Abbott,1978 **Advanced Physical Chemistry: Molecules, Structure, and Spectra** Jeff C. Davis (jr.),1965 **Advanced Physical Chemistry** Jeff C. Davis,2017-03-14 Advanced Physical Chemistry; Molecules, Structure, and Spectra Jefferson C. Davis,1967 A System of Physical Chemistry William Cudmore McCullagh Lewis,1916 **A System of Physical Chemistry** William Cudmore McCullagh Lewis,1922 Academic Press Dictionary of Science and Technology Christopher G. Morris,Academic Press,1992-08-27 A Dictionary of Science and Technology Color Illustration Section Symbols and Units Fundamental Physical Constants Measurement Conversion Periodic Table of the Elements Atomic Weights Particles The Solar System Geological Timetable Five Kingdom Classification of Organisms Chronology of Modern Science Photo Credits **A System of Physical Chemistry** William Cudmore MacCullagh Lewis,William C.Mcc Lewis,1918 **Molecular Physical Chemistry for Engineering Applications** Florin Emilian Daneş,Silvia Daneş,Valeria Petrescu,Eleonora-Mihaela Ungureanu,2021-07-06 This textbook introduces the molecular side of physical chemistry It offers students and practitioners a new approach to the subject by presenting numerous applications and solved problems that illustrate the concepts introduced for varied and

complex technical situations The book offers a balance between theory tools and practical applications The text aims to be a practical manual for solving engineering problems in industries where processes depend on the chemical composition and physical properties of matter The book is organized into three main topics I the molecular structure of matter II molecular models in thermodynamics and III transport phenomena and mechanisms Part I presents methods of analysis of the molecular behavior in a given system while the following parts use these methods to study the equilibrium states of a material system and to analyze the processes that can take place when the system is in a state of non equilibrium in particular the transport phenomena Molecular Physical Chemistry for Engineering Applications is designed for upper level undergraduate and graduate courses in physical chemistry for engineers applied physical chemistry transport phenomena colloidal chemistry and transport transfer processes The book will also be a valuable reference guide for engineers technicians and scientists working in industry Offers modeling techniques and tools for solving exercises and practical cases Provides solutions and conclusions so students can follow results more closely Step by step problem solving enables students to understand how to approach complex issues

World Directory of Crystallographers Allan L. Bednowitz, Armin P. Segmüller, 2013-04-17 A brief historical account of the background leading to the publication of the first four editions of the World Directory of Crystallographers was presented by G Boom in his preface to the Fourth Edition published late in 1971 That edition was produced by traditional typesetting methods from compilations of biographical data prepared by national Sub Editors The major effort required to produce a directory by manual methods provided the impetus to use computer techniques for the Fifth Edition The account of the production of the first computer assisted Directory was described by S C Abrahams in the preface of the Fifth Edition Computer composition which required a machine readable data base offered several major advantages The choice of typeface and range of characters was flexible Corrections and additions to the data base were rapid and once established it was hoped updating for future editions would be simple and inexpensive The data base was put to other Union uses such as preparation of mailing labels and formulation of lists of crystallographers with specified common fields of interest The Fifth Edition of the World Directory of Crystallographers was published in June of 1977 the Sixth in May of 1981 The Subject Indexes for the Fifth and Sixth Editions were printed in 1978 and 1981 respectively both having a limited distribution

A System of Physical Chemistry William Cudmore McCullagh Lewis, 1925
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