

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) indicated. Show the calculation for pH or percent dissociation when required. Write your answers in the space provided.

1. What substance is H^+ in water? Hydrogen ion is H^+ in water.
2. What substance is OH^- in water? Hydroxide ion is OH^- in water.
3. What substance is H_2O in water? Water is H_2O in water.
4. What substance is H^+ in water? Hydrogen ion is H^+ in water.
5. What substance is OH^- in water? Hydroxide ion is OH^- in water.
6. What substance is H_2O in water? Water is H_2O in water.
7. What substance is H^+ in water? Hydrogen ion is H^+ in water.
8. What substance is OH^- in water? Hydroxide ion is OH^- in water.
9. What substance is H_2O in water? Water is H_2O in water.

1. H^+
2. OH^-
3. H_2O
4. H^+
5. OH^-
6. H_2O
7. H^+
8. OH^-
9. H_2O

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Advisory Council on College Chemistry



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