

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many water molecules form a hydrogen bond with a neighboring water molecule at typical temperatures? (100°)
- ☐ 1
☒ 2
☐ 3
☐ 4

2. At equilibrium $H_2O(l) \rightleftharpoons H_2O(g)$, where would water molecules be located more frequently?
- ☒ liquid
☐ gas
☐ solid
☐ all

3. The condensed state closest to the normal one:
- ☐ that of the elements
☐ that of the compounds
☒ that of the mixtures

4. Pure hexamethylenediamine $H_2N(CH_2)_6NH_2$ is a solid at room temperature. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☐ the solid
☐ the liquid
☒ the gas

5. Water is the study of the water-miscible system. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☐ the solid
☐ the liquid
☒ the gas

6. Water is the study of the water-miscible system. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☐ the solid
☐ the liquid
☒ the gas

7. Hexamethylenediamine is a solid at room temperature. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☒ the solid
☐ the liquid
☐ the gas

8. Pure hexamethylenediamine $H_2N(CH_2)_6NH_2$ is a solid at room temperature. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☒ the solid
☐ the liquid
☐ the gas

9. At equilibrium $H_2O(l) \rightleftharpoons H_2O(g)$, where would water molecules be located more frequently?
- ☐ liquid
☐ gas
☒ solid
☐ all

10. Hexamethylenediamine is a solid at room temperature. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☐ the solid
☒ the liquid
☐ the gas

11. Hexamethylenediamine is a solid at room temperature. It is a polar molecule with hydrogen bonding. Condensed at what is it (normal)?
- ☐ the solid
☒ the liquid
☐ the gas

Multiple Choice Stoichiometry Test

John T. Moore, Richard H. Langley



Multiple Choice Stoichiometry Test:

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