



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acid	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acid
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

Principles of Ionization

Water is a neutral molecule in which the hydrogen and oxygen atoms are covalently bonded together. It is the ionization of these atoms that gives water its acidic properties.

Water ionizes to a very small extent to form a very small concentration of hydrogen ions (H^+) and hydroxide ions (OH^-). The concentration of hydrogen ions is 10^{-7} M.

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Ionization

<ul style="list-style-type: none">Water ionizationAcidStrong acidWeak acidNeutralBasicStrong baseWeak base	<ul style="list-style-type: none">Water ionizationAcidStrong acidWeak acidNeutralBasicStrong baseWeak base
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