

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate formula for each of the following acids, bases, and salts, showing dissociation reactions, and calculate the pH of each solution.

1. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$
2. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$
3. Hydrocyanic acid, HCN , $K_a = 4.0 \times 10^{-10}$
4. Hydrobromic acid, HBr , $K_a = 1.0 \times 10^3$
5. Hydrochloric acid, HCl , $K_a = 1.0 \times 10^7$
6. Hydroiodic acid, HI , $K_a = 1.0 \times 10^9$
7. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
8. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
9. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
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20. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$

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