

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?
 a. 1
 b. 10
 c. 100
 d. 1000

2. At temperature T_1 , 1 mol of H_2O vapor has kinetic energy E_1 . At temperature T_2 , 1 mol of H_2O vapor has kinetic energy E_2 .
 a. $E_2 = E_1$
 b. $E_2 = 2E_1$
 c. $E_2 = 4E_1$
 d. $E_2 = 8E_1$

3. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.
 a. 40.7 kJ/mol
 b. 46.7 kJ/mol
 c. 52.7 kJ/mol
 d. 58.7 kJ/mol

4. First Law of Thermodynamics: $\Delta U = Q + W$. If $Q = 100 \text{ J}$ and $W = 50 \text{ J}$, then ΔU is:
 a. 150 J
 b. 100 J
 c. 50 J
 d. 0 J

5. Water is the most abundant molecule in the universe. It is a polar molecule. It is a liquid at room temperature. It is a solid at low temperature. It is a gas at high temperature.
 a. liquid
 b. solid
 c. gas
 d. plasma

6. Water is the most abundant molecule in the universe. It is a polar molecule. It is a liquid at room temperature. It is a solid at low temperature. It is a gas at high temperature.
 a. 100 J
 b. 150 J
 c. 200 J
 d. 250 J

7. A sample of water is heated from 10°C to 20°C. The heat added is 100 J. The heat capacity of water is 4.18 J/(mol·K). The number of moles of water is:
 a. 100 J
 b. 200 J
 c. 300 J
 d. 400 J

8. First Law of Thermodynamics: $\Delta U = Q + W$. If $Q = 100 \text{ J}$ and $W = 50 \text{ J}$, then ΔU is:
 a. 150 J
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 c. 50 J
 d. 0 J

9. At temperature T_1 , 1 mol of H_2O vapor has kinetic energy E_1 . At temperature T_2 , 1 mol of H_2O vapor has kinetic energy E_2 .
 a. $E_2 = E_1$
 b. $E_2 = 2E_1$
 c. $E_2 = 4E_1$
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10. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.
 a. 40.7 kJ/mol
 b. 46.7 kJ/mol
 c. 52.7 kJ/mol
 d. 58.7 kJ/mol

11. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.

Multiple Choice Stoichiometry Test

Richard H. Langley, John Moore



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