

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HBr~~ **hydrobromic acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~H<sub>2</sub>SO<sub>4</sub>~~ **sulfuric acid**
- ~~HClO<sub>4</sub>~~ **perchloric acid**
- ~~Sodium hydroxide~~ **NaOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~barium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~ammonium hydroxide~~ **ammonia**

~~HClO<sub>3</sub> is a STRONG acid & it does not completely dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

**HCl**

~~HBr~~

~~HNO<sub>3</sub>~~

~~Mg(OH)<sub>2</sub>~~

**NaOH**

**NaOH**

~~NaOH~~

**NaOH**

3. What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**

4.  $\text{BaCu}_2\text{O}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  **$\text{BaCu}_2\text{O}_4 + \text{H}_2\text{O} \rightarrow \text{BaCu}_2\text{O}_3 + \text{H}^+$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{BaCu}_2\text{O}_4$  is a weak acid.**
- Would you define  $\text{BaCu}_2\text{O}_4$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{BaCu}_2\text{O}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{BaCu}_2\text{O}_4$ :  **$\text{BaCu}_2\text{O}_3$**
- Identify  $\text{BaCu}_2\text{O}_4$  as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5.  $\text{F}^-$  will react with water and gain a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{F}^-$  is a weak base.**
- Would you define  $\text{F}^-$  as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{F}^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of  $\text{F}^-$ : **HF**
- Identify  $\text{F}^-$  as polyprotic, amphoteric, or neither. **neither**

6.  $\text{HPO}_4^{2-}$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  **$\text{HPO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{PO}_4^- + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{HPO}_4^{2-}$  is a weak acid.**
- Would you define  $\text{HPO}_4^{2-}$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{HPO}_4^{2-}$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{HPO}_4^{2-}$ :  **$\text{PO}_4^{3-}$**
- Identify  $\text{HPO}_4^{2-}$  as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and can take a proton) and it has H's (and can lose one).**

# 19 3 Acids And Bases Answer Key Chemistry

**Simeon Shaw**



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web jul 27 2023 apj abdul kalam death anniversary is celebrated on 27th july every year around the world who is also known as missile man of india 27 th july 2023 is celebrated as 8 th death anniversary of apj abdul kalam