

Biology Review Sheet – 2nd Semester 2011

My final exam date is _____

TAXONOMY (Ch. 17)

1. Know that Carolus Linnaeus came up with the basis of our current taxonomy system and Aristotle had a flawed taxonomy system in 350 BCE
2. Know that the kingdoms go from relatively simple (eubacteria) to more complex (animals)
3. Know the 8 levels in the classification Hierarchy (D, K, P, C, O, F, G, S)
4. Know how to correctly write a scientific name and use Binomial Nomenclature
5. Know how to use a Dichotomous Key
6. Know how to interpret a Cladogram
7. Know how to read a Phylogenetic Tree

EVOLUTION (CH. 14 and Ch. 15)

1. Know who Charles Darwin is and his Theory of Natural Selection
2. Know who Jean Baptiste Lamarck is and his incorrect ideas of "Law of Use and Disuse" and "Law of Acquired Traits"
3. Know directional, disruptive, and stabilizing selection
4. Know the types of evidence for evolution (embryology, biochemistry, fossils, homologous, analogous and vestigial structures)
5. Know types / mechanisms of evolution (adaptive radiation, convergent evolution, co-evolution, sexual selection, gradualism, punctuated equilibrium)
6. Know who Stanley Miller and Harold Urey were and their experiment that determined how simple organic molecules could be formed from atmospheric gasses

BACTERIA-VIRUS - IMMUNITY UNIT (Ch.18 and Ch. 37)

1. Identify structures and parts of a virus
2. Know the characteristics of a virus and why they are ultimately considered to be non-living by science
3. Understand the difference between the Lytic Cycle and the Lysogenic Cycle
4. Identify the different shapes and groupings of bacteria
5. Understand the similarities and differences between the Kingdom Eubacteria and the Kingdom Archaeabacteria (extremophiles)
6. Know the sexual and asexual reproductive methods of eubacteria
7. Know the difference between Active and Passive Immunity
8. Know the major body organs and structures involved in the Immune System
9. Know the difference between antiseptic and disinfectant

2nd Semester Exam Review Sheet Answer

Barbara Fine Clouse

2nd Semester Exam Review Sheet Answer:

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3 655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1 1 we could experimentally determine the number of moles of the acid in the solution to be 000731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the k a or k b from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate k a1 and k a2 for a polyprotic acid by the end of this lab students should be able to

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq NaOH aq $\text{Na}_2\text{C}_2\text{O}_4$ $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
naoh aq c6h4 cooh cook aq c6h4 coona cook aq h2o l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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