

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element within each molecule or formula.

1. Hydrochloric acid, HCl (Equation 19.1)

2. Acetic acid, CH_3COOH (Equation 19.1)

3. Nitric acid, HNO_3 (Equation 19.1)

4. Sulfuric acid, H_2SO_4 (Equation 19.1)

5. Phosphoric acid, H_3PO_4 (Equation 19.1)

6. Hydrofluoric acid, HF (Equation 19.1)

7. Hydrobromic acid, HBr (Equation 19.1)

8. Hydroiodic acid, HI (Equation 19.1)

9. Hydrocyanic acid, HCN (Equation 19.1)

10. Hydroperfluoroboric acid, HF_2 (Equation 19.1)

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