

AP CHEMISTRY FINAL EXAM REVIEW 2

1. Is energy absorbed or released in an endothermic reaction?: Absorbed.
2. What is the sign of delta H for an exothermic reaction?: Exothermic reactions release energy. The sign of delta H is negative.
3. When you touch a flask and it feels hot, is this an endothermic or exothermic process?: Exothermic; heat is being given off.
4. Define specific heat capacity.: The amount of heat required to raise 1g of a substance to 1 degrees Celsius.
5. $q=mc(\Delta T)$: What does "q" stand for?: Heat (kJ)
6. $q=mc(\Delta T)$: What does "c" stand for?: Specific heat capacity (J / g degreesC)
7. What is a hydrogen ion donor called?: An acid.
8. What is a hydrogen ion acceptor called?: A base.
9. Feels slippery; acid or base?: Base.
10. Tastes sour; acid or base?: Acid.
11. Tastes bitter; acid or base?: Base.
12. $pH < 7$: acid, base, or neutral?: Acid
13. $pH > 7$: acid, base, or neutral?: Base
14. Define neutralization.: When an acid reacts with a base to form water and a salt. $pH = 7$.
15. Soap; acid or base?: Base (tastes bitter)
16. Ammonia, acid or base?: Base (drain cleaners are bases).
17. Lemonade; acid or base?: Acid (tastes sour)
18. How many valence electrons are in an atom of chlorine?: 7
19. How many valence electrons does aluminum have?: 3
20. Define Reaction Rate: The decrease in concentration of reactants or the increase in concentration of products with time.
21. How does increasing temperature affect the rate of reaction?: Increasing temperature causes the kinetic energy to increase. Thus, molecules are moving faster and colliding more frequently. This causes the rate of reaction to increase.
22. If I decrease the concentration of reactant, what happens to the rate of reaction?: Decreasing concentration leads to less reactants available for collision. Therefore, reaction rate decreases.
23. How does surface area influence reaction rate?: By increasing the surface area of reactant molecules, we are increasing the space or point of contact that these molecules have available

2nd Semester Chemistry Final Exam Review Answers

Saul Wischnitzer, Edith Wischnitzer

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Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky Andrew M. Olney, Irene-Angelica

Chounta, Zitao Liu, Olga C. Santos, Ig Ibert Bittencourt, 2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8-12 2024. The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions. They are organized in topical sections as follows: Part One: Blue Sky, Industry, Innovation and Practitioner; Part Two: Late Breaking Results, Doctoral Consortium, Workshops and Tutorials.

Quarterly Review of Distance Education Michael Simonson, Charles Schlosser, 2018-10-01 The Quarterly Review of Distance Education is a rigorously refereed journal publishing articles, research briefs, reviews, and editorials dealing with the theories, research, and practices of distance education. The Quarterly Review publishes articles that utilize various methodologies that permit generalizable results which help guide the practice of the field of distance education in the public and private sectors. The Quarterly Review publishes full length manuscripts as well as research briefs, editorials, reviews of programs and scholarly works and columns. The Quarterly Review defines distance education as institutionally based formal education in which the learning group is separated and interactive technologies are used to unite the learning group.

Resources in Education, 1990-05 **Essentials of Physical Chemistry** Don Shillady, 2011-07-27 At a time when U.S. high school students are producing low scores in mathematics and science on international examinations, a thorough grounding in physical chemistry should not be considered optional for science undergraduates. Based on the author's thirty years of teaching, *Essentials of Physical Chemistry* merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner. Reflecting the latest ACS guidelines, the book can be used as a one- or two-semester course and includes special topics suitable for senior projects. The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students' abilities. It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules. Emphasizing spectroscopy, the text is built on a foundation of quantum chemistry and more mathematical detail and examples. It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts. Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations. Mini biographies of famous scientists make the presentation more interesting from a people point of view. Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique. Covering key topics such as the critical point of a van der Waals gas, the Michaelis-Menten

equation and the entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry

Reports from the Classroom Sarah H. Huyvaert,1995 **World of Chemistry** Melvin D. Joesten,2004 **Nursing School Entrance Exam Review** William H. Brown, Frederick A. Bettelheim, Jerry March,2004-04 **5 Steps to a 5 AP Chemistry, 2012-2013 Edition** Richard H. Langley, John Moore, 2011-06-10

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Biology Edward C. Gruber, 1963 **Chemical Engineering Education** ,1989 **Lorena** Dennison Ambrose, 2007-04-09 Lorena an unworldly young girl who has always bowed to her

father's tyrannical demands meets a charming young man who is quite taken with her. When they are seen together word gets back to her father and he orders Lorena to stay away from the young man who is the son of his enemy. The enmity arose when he lost an expensive court case against him. He breaks up the budding romance by locking Lorena in her room and she is a prisoner in her parents' house. She is forced to marry a much older man and becomes a prisoner on her husband's ranch. He abuses her physically and psychologically. Read how she escapes and rebuilds her life.

Journal of Developmental Education, 2002 *Foundations of Learning Activities Book* Karl Krumsieg, Marie Baehr, 2000 The companion to Foundations of Learning for instructors who want an active learning component in their course. More than 40 activities in a variety of formats case studies guided discovery group discussion problem solving etc put students in situations where they practice the concepts and learning skills presented in Foundations of Learning.

Iterations, II Russell Batt, John W. Moore, 1987 **Paperbound Books in Print**, 1966 **Respiratory Care**, 2001

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