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Kenya

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp v = 1 n = 1 v = 2 n = 0.0896 l n = 22.4 l 1 mole n = 0.0040 moles 12 experimental molar mass g/mole of butane 0.23 g 0.0040 moles 58 g/mole 13 the molecular formula of butane is C4H10

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration

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