

# AP Calculus Notes:

AB:

Pre-calc Stuff:

- Know all of this (it's vital): [\(insert quizlet link\)](#)

Limits:

- If the limit of a function as  $x$  approaches a constant ( $c$ ) exists then the function is interrupted.
- Types of Discontinuities:
  - Hole (removable)
  - Jump Discontinuity (non-removable)
  - Infinite Discontinuity (non-removable)
- If limit from the right isn't the same as the limit from the left then the limit does not exist (DNE).
- Limit of a constant is the constant itself.
- Plug in the value for  $x$  ( $c$ ) to solve for basic limits (duh).
- To solve limit questions where you get the indeterminate form of  $0/0$  or  $\infty/\infty$  you can:
  - SUBSTITUTE (ALWAYS CONSIDER THIS FIRST! But since you got the indeterminate form move on to the next steps)
  - Factor
  - Rationalize
  - Simplify/Use Identities
  - Common denominator (usually for numerator when there are fractions).

# 2008 Ap Calculus Ab Practice Exam Multiple Choice Answers

**Rose Kearney Nunnery**



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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{H}_2\text{SO}_4$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$   $\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O}$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{OH}^{-} + 2\text{e}^{-} \rightarrow 4\text{HCN} + 2\text{OH}^{-}$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O}$  the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh<sup>-</sup> ions or the h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo<sup>-</sup> and cr oh<sub>4</sub><sup>-</sup> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh<sup>-</sup> ions to form bromide ions and bro<sup>-</sup> ions second the bro<sup>-</sup> ions react to form bromide ions and bromate v ions

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web mar 13 2023 balancing redox equations by the ion electron method there are two principal methods for balancing redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h<sup>+</sup> ions or the h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> and chloride cl<sup>-</sup> ions in acidic solution created by jay

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web balance each half reaction in basic solution  $4\text{S} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO}_3^- + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox



reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2$   $\text{Cr}_2\text{O}_7^{2-}$  9  $\text{TeO}_3$   $\text{N}_2\text{O}_4$   $\text{O}_2$   $\text{Cr}_3$  10  $\text{ReO}_4$  11  $\text{PbO}_2$  12  $\text{I}_2$   $\text{Pb}_2$  12 as  $\text{NO}_3$   $\text{IO}_3$  re  $\text{IO}_3$   $\text{H}_2\text{AsO}_4$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$   $\text{OH}^-$  and  $\text{e}^-$  in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{HNO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$   $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$   $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$   $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and