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Amanda Lamb



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$$\begin{array}{l} \text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3 \\ \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2 \\ \text{NaCl} + \text{F}_2 \rightarrow 2 \text{NaF} + \text{Cl}_2 \\ 2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O} \\ \text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2 \\ \text{Al} + \text{Br}_3 \rightarrow \text{AlBr}_3 \\ \text{K}_2\text{SO}_4 + \text{KBr} + \text{Al}_2\text{SO}_4 + \text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{C}_3\text{H}_8 + \text{O}_2 \end{array}$$

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4$
 $2\text{O}_2 \rightarrow 1\text{CO}_2 + 2\text{H}_2\text{O}$

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 $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$
 $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$
 $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

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 $3\text{Fe} + 2\text{O}_3 \rightarrow 3\text{FeO}$
 $2\text{Fe} + 3\text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + 3\text{C}$
 $3\text{BaO} + \text{Ba} \rightarrow 4\text{Ba}$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $4\text{Mg} + \text{NH}_4\text{PO}_4 \rightarrow 4\text{Mg}_3\text{P}_2 + 7\text{N}_2 + 3\text{H}_2 + \text{O}_2$
 $4\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{O}_3 + 3\text{H}_2 + 2\text{H}_2\text{O}$
 $4\text{P} + 10\text{HCl} \rightarrow 3\text{H}_3\text{PO}_3 + 4\text{H}_2 + 3\text{O}_2$
 $4\text{HNO}_2 \rightarrow 2\text{H}_2\text{O} + 4\text{HNO}$

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web mar 14 2023 the key to balancing chemical equations the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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