

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH of a given solution.
3. Calculate the pH of a weak acid solution.

Vocabulary

- | | | |
|-------------------|-------------------------------|---------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. basic solution | 6. neutral solution |
| | | 7. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or hydroxide ion concentration of the solution described. Round the concentration to the correct significant figures. Express acids and bases in scientific notation.

1. What substance is _____, contains hydrogen ions?
2. _____ is the concentration of those acids present in water and equal to _____.
3. _____ is the concentration of those acids present in water and equal to _____.
4. _____ is the concentration of those acids present in water and equal to _____.
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18. _____ is the concentration of those acids present in water and equal to _____.
19. _____ is the concentration of those acids present in water and equal to _____.
20. _____ is the concentration of those acids present in water and equal to _____.

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