

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|------------------------------------|
| acid | acid dissociation constant (K_a) | acid ionization constant (K_a) |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the K_a expression for each of the following weak acids, and calculate the K_a value for each acid. (Use the K_a values in the table below.)

1. Acetic acid, CH_3COOH
2. Hydrofluoric acid, HF
3. Hydrocyanic acid, HCN
4. Hydrobromic acid, HBr
5. Hydroiodic acid, HI
6. Hydrochloric acid, HCl
7. Hydroperchloric acid, $HClO_4$
8. Hydroperbromic acid, $HBrO_4$
9. Hydroperchloric acid, $HClO_4$
10. Hydroperbromic acid, $HBrO_4$
11. Hydroperchloric acid, $HClO_4$
12. Hydroperbromic acid, $HBrO_4$
13. Hydroperchloric acid, $HClO_4$
14. Hydroperbromic acid, $HBrO_4$
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16. Hydroperbromic acid, $HBrO_4$
17. Hydroperchloric acid, $HClO_4$
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19. Hydroperchloric acid, $HClO_4$
20. Hydroperbromic acid, $HBrO_4$

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