

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_m$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $v_f^*$  (so for most molecules at normal temperatures  $v_f^* = 2$  for linear molecules, and 3 for non-linear molecules), then

# Physical Chemistry 9th Edition Instructors Solution Guide

**Lei Shi**

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*Instructor's Solutions Manual to Accompany Atkins' Physical Chemistry, Ninth Edition* C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 The Instructor's solutions manual to accompany Atkins' Physical Chemistry provides detailed solutions to the b exercises and the even numbered discussion questions and problems that feature in the ninth edition of Atkins' Physical Chemistry. The manual is intended for instructors and consists of material that is not available to undergraduates. The manual is free to all adopters of the main text. **Solutions to Exercises** Roxy Wilson, 2002-05 Full solutions to all end of chapter exercises in the text are provided. With an instructor's permission this manual may be made available to students.

**Children's Books in Print, 2007**, 2006 *Subject Guide to Books in Print*, 2001 *Educational Times*, 1884

**Instructor's Resource Manual for Woolfolk, Educational Psychology, Ninth Edition** Angela M. O'Donnell, 2004

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