

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Write the appropriate words in the blanks to complete the following sentences.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

1. 1.0×10^{-4}
2. 1.0×10^{-10}
3. 1.0×10^{-4}
4. 1.0×10^{-10}
5. 1.0×10^{-4}
6. 1.0×10^{-10}
7. 1.0×10^{-4}
8. 1.0×10^{-10}
9. 1.0×10^{-4}
10. 1.0×10^{-10}

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Padhraic Smyth



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