



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log \frac{1}{[\text{H}^+]}$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

### Principles of Ionization

Acid dissociation is the process of ionization in which an acid dissociates into ions. This process is reversible.

Water dissociation is the process of ionization in which water dissociates into ions. This process is reversible.

Strong acids dissociate completely in water. Strong acids are strong electrolytes.

Weak acids dissociate partially in water. Weak acids are weak electrolytes.

### Acid Dissociation

<ul style="list-style-type: none"><li>strong acid</li><li>weak acid</li><li>strong base</li><li>weak base</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>strong acid</li><li>weak acid</li><li>strong base</li><li>weak base</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>strong acid</li><li>weak acid</li><li>strong base</li><li>weak base</li><li>conjugate acid</li><li>conjugate base</li></ul>
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