

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HBr~~ **hydrobromic acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~H<sub>2</sub>SO<sub>4</sub>~~ **sulfuric acid**
- ~~HClO<sub>4</sub>~~ **perchloric acid**
- ~~Sodium hydroxide~~ **NaOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~barium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~ammonium hydroxide~~ **NaOH**

~~HClO<sub>3</sub> for chloric acid & HNO<sub>2</sub> for nitrous acid. Dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

**HCl**

~~HBr~~

~~HNO<sub>3</sub>~~

~~Mg(OH)<sub>2</sub>~~

**NaOH**

**NaOH**

~~NaOH~~

**NaOH**

3. What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**

4.  $\text{H}_2\text{C}_2\text{O}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HC}_2\text{O}_4^- + \text{H}_3\text{O}^+$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{H}_2\text{C}_2\text{O}_4$  is a weak acid.**
- Would you define  $\text{H}_2\text{C}_2\text{O}_4$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{H}_2\text{C}_2\text{O}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{H}_2\text{C}_2\text{O}_4$ :  **$\text{HC}_2\text{O}_4^-$**
- Identify  $\text{H}_2\text{C}_2\text{O}_4$  as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5.  $\text{F}^-$  will react with water and gain a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{F}^-$  is a weak base.**
- Would you define  $\text{F}^-$  as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{F}^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of  $\text{F}^-$ : **HF**
- Identify  $\text{F}^-$  as polyprotic, amphoteric, or neither. **neither**

6.  $\text{HSO}_4^-$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons \text{SO}_4^{2-} + \text{H}_3\text{O}^+$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{HSO}_4^-$  is a weak acid.**
- Would you define  $\text{HSO}_4^-$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{HSO}_4^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{HSO}_4^-$ :  **$\text{SO}_4^{2-}$**
- Identify  $\text{HSO}_4^-$  as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will take a proton) and it has H's (and will lose one too).**

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**Charles Loudon Bloxam**



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