

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, from  $pK_a$  and  $pK_b$  values.
2. Calculate  $pH$  and  $pOH$  for weak acids and bases, and vice versa, from  $K_a$  and  $K_b$  values.
3. Calculate the percent ionization of weak acids.

## Vocabulary

- |                            |                             |                             |
|----------------------------|-----------------------------|-----------------------------|
| acid dissociation constant | 1. conjugate acid           | 2. acid ionization constant |
| acid ionization constant   | 3. base                     |                             |
|                            | 4. base ionization constant |                             |

## Key Equations

- |                                       |                                      |
|---------------------------------------|--------------------------------------|
| 1. $K_a = 10^{-pK_a}$ (Equation 19.1) | 2. $pH = -\log[H^+]$ (Equation 19.2) |
| 3. $pOH = 14 - pH$ (Equation 19.3)    |                                      |

## Part A Completion

Write the appropriate number in the blank space corresponding to the correct answer. Write the number of the correct answer in the blank space after the question.

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
2. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
3. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
4. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
5. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
6. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
7. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
8. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.

9. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
10. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
11. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
12. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
13. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
14. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.
15. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $pH$  of a 0.1 M solution of this acid is \_\_\_\_\_.
16. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $pOH$  of a 0.1 M solution of this base is \_\_\_\_\_.

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**Sabine Zange**



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