

CHEM 12: Hess's Law Worksheet

1. Calculate the ΔH for the reaction: $\text{C}_2\text{H}_6(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$, from the following data:



$$\Delta H = -1411 \text{ kJ / mole}$$



$$\Delta H = -1560 \text{ kJ / mole}$$



$$\Delta H = -285.8 \text{ kJ / mole}$$

2. Calculate the ΔH for the reaction:



from the following data:



$$\Delta H = -180.5 \text{ kJ / mole}$$



$$\Delta H = -91.8 \text{ kJ / mole}$$



$$\Delta H = -483.6 \text{ kJ / mole}$$

3. Find the ΔH_f° (heat of formation) for acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, using the following thermochemical data:



$$\Delta H = -875 \text{ kJ / mole}$$



$$\Delta H = -394.51 \text{ kJ / mole}$$



$$\Delta H = -285.8 \text{ kJ / mole}$$

Pure Factor C(s) is in the form of graphite.

Acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, actually has the structure: CH_3COOH



4. Calculate the ΔH for the reaction: $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + 3\text{H}_2(\text{g})$, from the following data:



$$\Delta H = -91.8 \text{ kJ / mole}$$



$$\Delta H = -74.9 \text{ kJ / mole}$$



$$\Delta H = +270.3 \text{ kJ / mole}$$

12 3 Hess S Law Answers

Brian F. Woodfield,Matthew C. Asplunc

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Thermodynamics Ricardo Morales-Rodriguez, 2012-10-03 This book presents the selection of various high level contributions involving thermodynamics. The book goes from the fundamentals up to several applications in different scientific fields. The content of the book has been classified in six sections: Classical Thermodynamics, Statistical Thermodynamics, Property Prediction in Thermodynamics, Material and Products Non Equilibrium and Thermodynamics in Diverse Areas. The classification of the book aims to provide to the reader the facility of finding the desired topic included in the book. It is expected that this collection of chapters will contribute to the state of the art in the thermodynamics area.

Elements of Physical Chemistry Peter William Atkins, Julio De Paula, 2017 This revision of the introductory textbook of physical chemistry has been designed to broaden its appeal particularly to students with an interest in biological applications.

Solution and Electrolysis Sir William Cecil Dampier, 1895 Physical Chemistry Peter Atkins, Julio de Paula, Ronald Friedman, 2014 This title takes an innovative molecular approach to the teaching of physical chemistry. The authors present the subject in a rigorous but accessible manner allowing students to gain a thorough understanding of physical chemistry.

Prentice Hall Chemistry, 2000 **Solution & Electrolysis** William Cecil Dampier Whetham, 1895

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Physical Chemistry for the Biosciences, second edition Raymond Chang, Charles M. Lovett, 2025-02-25 Physical Chemistry for the Biosciences has been optimized for a one semester course in physical chemistry for students of biosciences or a course in biophysical chemistry. Most students enrolled in this course have taken general chemistry, organic chemistry, and a year of physics and calculus. Fondly known as Baby Chang, this best selling text is back in an updated second edition for the one semester physical chemistry course. Carefully crafted to match the needs and interests of students majoring in the life sciences, Physical Chemistry for the Biosciences has been revised to provide students with a sophisticated appreciation for physical chemistry as the basis for a variety of interesting biological phenomena. Major changes to the new edition include: Discussion of intermolecular forces in chapter 1, Detailed discussion of protein and nucleic acid structure providing students with the background needed to fully understand the biological applications of thermodynamics and kinetics described later in the book, Expanded and updated descriptions of biological examples such as protein misfolding diseases, photosynthesis and vision.

Competition Science Vision, 2008-02 Competition Science Vision monthly magazine is published by Pratiyogita

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Instructor's Manual Brian F. Woodfield,Matthew C. Asplunc,2006 NEW

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