

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, from pK_a and pK_b values.
2. Calculate pH and pOH for weak acids and bases, and vice versa, from K_a and K_b values.
3. Calculate the percent ionization of weak acids.

Vocabulary

- | | | |
|----------------------------|----------------------------------|-----------------------------|
| acid dissociation constant | 1. conjugate base of a weak acid | 3. acid ionization constant |
| acid ionization constant | 2. conjugate acid of a weak base | 4. pH |

Key Equations

- | | |
|---------------------------------------|--|
| 1. $K_a = 10^{-pK_a}$ (Equation 19.1) | 2. $pH = -\log[H^+]$ (Equation 19.2) |
| 3. $pH = 14 - pOH$ (Equation 19.3) | 4. $pOH = -\log[OH^-]$ (Equation 19.4) |

Part A Completion

Write the appropriate number in each blank space, and write clearly and correctly the names of the acids or bases represented within each blank space, or solution.

1. Acetic acid has a K_a of 1.8×10^{-5} . Its pK_a is _____.
2. Hydrofluoric acid has a pK_a of 3.17. Its K_a is _____.
3. A solution of hydrofluoric acid has a pH of 2.11. The concentration of H^+ in this solution is _____ M .
4. A solution of hydrofluoric acid has a pH of 2.11. The concentration of HF in this solution is _____ M .
5. A solution of hydrofluoric acid has a pH of 2.11. The concentration of F^- in this solution is _____ M .
6. A solution of hydrofluoric acid has a pH of 2.11. The concentration of OH^- in this solution is _____ M .
7. A solution of hydrofluoric acid has a pH of 2.11. The concentration of H_2O in this solution is _____ M .
8. A solution of hydrofluoric acid has a pH of 2.11. The concentration of H_2O in this solution is _____ M .

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Marcel A. Müller



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