

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                             |                 |
|-------------------|-----------------------------|-----------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid    |
| 4. autoionization | 5. $\text{pH}$              | 6. $\text{pOH}$ |

## Key Equation

- |                                     |                                  |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Read the statement and fill in the blank with the word or phrase that is indicated in parentheses. Write the word or phrase in the space provided.

1. When a substance ionizes, it releases hydrogen ions. (acidic)
2. The concentration of  $\text{H}^+$  in the autoionization of water is  $1.0 \times 10^{-7}$  moles per liter at 25°C. (neutral)
3. The  $\text{pH}$  of a solution is defined as the negative logarithm of the concentration of  $\text{H}^+$ . (acidic)
4. The  $\text{pH}$  of a solution is defined as the negative logarithm of the concentration of  $\text{H}^+$ . (acidic)
5. The  $\text{pH}$  of a solution is defined as the negative logarithm of the concentration of  $\text{H}^+$ . (acidic)
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# 19 2 Hydrogen Ions And Acidity Answer Key

**Valerie C. Scanlon, Tina Sanders**



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