



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base in terms of the Arrhenius theory.
- Define strong and weak acids.

### Definitions

- Hydrogen ion:  $H^+$
- Hydroxide ion:  $OH^-$
- Acid: a substance that dissociates in water to produce hydrogen ions.
- Base: a substance that dissociates in water to produce hydroxide ions.
- Neutral: a substance that dissociates in water to produce equal numbers of hydrogen and hydroxide ions.

### Key Equations

- $pH = -\log[H^+]$
- $pOH = -\log[OH^-]$

$$pH + pOH = 14$$

### Acids & Bases

Acids and bases are substances that dissociate in water to produce hydrogen and hydroxide ions, respectively.

Strong acids dissociate completely in water to produce hydrogen ions. Examples of strong acids include sulfuric acid, hydrochloric acid, and nitric acid.

Weak acids dissociate partially in water to produce hydrogen ions. Examples of weak acids include acetic acid, carbonic acid, and formic acid.

Strong bases dissociate completely in water to produce hydroxide ions. Examples of strong bases include sodium hydroxide, potassium hydroxide, and calcium hydroxide.

Weak bases dissociate partially in water to produce hydroxide ions.

### Acid-Base Reactions

- Neutralization: a reaction between an acid and a base that produces a salt and water.
- Acid-base indicators: substances that change color when they react with acids or bases.
- Titration: a method for determining the concentration of an acid or base by titrating it with a standard solution of a strong acid or base.
- Titration curve: a graph showing the pH of a solution as it is titrated with a strong acid or base.
- Half-equivalence point: the point on a titration curve where the pH is equal to 7.
- Equivalence point: the point on a titration curve where the pH is at its maximum or minimum.
- Post-equivalence point: the point on a titration curve where the pH is greater than 7.
- Titration endpoint: the point on a titration curve where the color of the indicator changes.

## 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Bill Scott**

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