



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base in terms of the Arrhenius theory.
- Define strong and weak acids.

### Definitions

- Hydrogen ion
- Hydrogen ion concentration
- Acid
- Base
- Strong acid
- Weak acid
- Strong base
- Weak base

### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

### Principles of Acids and Bases

Acids and bases are defined in terms of their ability to ionize in water. Acids are proton donors, and bases are proton acceptors.

Strong acids dissociate completely in water. The concentration of these acids is equal to the concentration of hydrogen ions.

Weak acids dissociate partially in water. The concentration of hydrogen ions is less than the concentration of the acid.

Strong bases dissociate completely in water. The concentration of hydroxide ions is equal to the concentration of the base.

Weak bases dissociate partially in water. The concentration of hydroxide ions is less than the concentration of the base.

### Acid-Base Properties

- Acidic
- Neutral
- Basic
- Alkaline

## 19 2 Hydrogen Ions And Acidity Answer Key

**JL Elias**

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