

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element within each molecule or formula.

1. Hydrochloric acid,  $HCl$  (Equation 19.1)

2. Acetic acid,  $CH_3COOH$  (Equation 19.1)

3. Nitric acid,  $HNO_3$  (Equation 19.1)

4. Sulfuric acid,  $H_2SO_4$  (Equation 19.1)

5. Phosphoric acid,  $H_3PO_4$  (Equation 19.1)

6. Hydrofluoric acid,  $HF$  (Equation 19.1)

7. Hydrobromic acid,  $HBr$  (Equation 19.1)

8. Hydroiodic acid,  $HI$  (Equation 19.1)

9. Hydrocyanic acid,  $HCN$  (Equation 19.1)

10. Hydroperfluoroboric acid,  $HFBO_2$  (Equation 19.1)

11. Hydroperfluoroboric acid,  $HFBO_2$  (Equation 19.1)

1. Hydrochloric acid
2. Acetic acid
3. Nitric acid
4. Sulfuric acid
5. Phosphoric acid
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10. Hydroperfluoroboric acid
11. Hydroperfluoroboric acid

# 19 2 Hydrogen Ions And Acidity Section Review

**Kevin T. Patton**



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