

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- k. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- l. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H<sub>2</sub>SO<sub>4</sub>**.

4. **HClO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{ClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{ClO}_4 + \text{H}_3\text{O}^+$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{HClO}_4$  in a weak acid.

c. Dissolve **1.00\text{M}**  $\text{NaClO}_4$  in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}**  $\text{HClO}_4$  in a portion of water.

d. In your **reaction** prediction, is  $\text{HClO}_4$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because  $\text{HClO}_4$  is a **strong** acid.

e. Identify the conjugate base of  $\text{HClO}_4$ :  **$\text{ClO}_4^-$**

f. Identify  $\text{NaClO}_4$  as **strong**, **weak**, or **none**: **strong**. Because it contains a conjugate base  $\text{ClO}_4^-$ .

5.  $\text{H}^+$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_2\text{O} + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{OH}^-$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{H}_2\text{O}$  in a weak base.

c. Dissolve **1.00\text{M}**  $\text{NaOH}$  in a portion of water and a portion of **1.00\text{M}** **NaCl**. Dissolve **1.00\text{M}**  $\text{H}_2\text{O}$  in a portion of water.

d. In your **reaction** prediction, is  $\text{H}_2\text{O}$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry base**. Because  $\text{H}_2\text{O}$  is a **strong** base.

e. Identify the conjugate acid of  $\text{H}_2\text{O}$ :  **$\text{H}_3\text{O}^+$**

f. Identify  $\text{NaOH}$  as **strong**, **weak**, or **none**: **strong**.

6. **H<sub>2</sub>SO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{SO}_4^+ + \text{H}_3\text{O}^+$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{H}_2\text{SO}_4$  in a weak acid.

c. Dissolve **1.00\text{M}**  $\text{NaHSO}_4$  in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}**  $\text{H}_2\text{SO}_4$  in a portion of water.

d. In your **reaction** prediction, is  $\text{H}_2\text{SO}_4$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because  $\text{H}_2\text{SO}_4$  is a **strong** acid.

e. Identify the conjugate base of  $\text{H}_2\text{SO}_4$ :  **$\text{HSO}_4^-$**

f. Identify  $\text{NaHSO}_4$  as **strong**, **weak**, or **none**: **strong**. Because it contains a conjugate base  $\text{HSO}_4^-$  (strong base) and a strong acid  $\text{H}_2\text{SO}_4$ .

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## **Table of Contents 19 3 Acids And Bases Answer Key Chemistry**

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
  - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry

- Personalized Recommendations
- 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
- 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists

5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks

- 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
- 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
- 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options

6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats

- ePUB, PDF, MOBI, and More
- 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
- 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
- Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
- Interactive Elements 19 3 Acids And Bases Answer Key Chemistry

8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry

9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry

- Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry

- Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
- Distinguishing Credible Sources

#### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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