

Multiple Choice

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?  
 a. 1  
 b. 10  
 c. 100  
 d. 1000

2. At standard  $T_p$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  vapor has higher enthalpy than liquid water.  
 a. True  
 b. False  
 c. Not sure  
 d. None

3. The enthalpy of fusion is the lowest for:  
 a. Water at 0°C  
 b. Ice at 0°C  
 c. Liquid water at 0°C  
 d. Solid water at 0°C

4. For the reaction  $\text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$ , the  $\Delta H$  for the reaction is  $-286 \text{ kJ/mol}$  of  $\text{H}_2$  consumed. The  $\Delta H$  for the reaction is  $-286 \text{ kJ/mol}$  of  $\text{H}_2$  consumed.  
 a. 286 kJ/mol  
 b. 572 kJ/mol  
 c. 143 kJ/mol  
 d. 286 kJ/mol

5. Water is the most abundant molecule in the atmosphere.  
 a. True  
 b. False  
 c. Not sure  
 d. None

6. Water is the most abundant molecule in the atmosphere.  
 a. 100%  
 b. 100%  
 c. 100%  
 d. 100%

7. A substance's specific heat is  $1 \text{ J/g}^\circ\text{C}$  at the boiling point. It is  $1 \text{ J/g}^\circ\text{C}$  at the boiling point.  
 a. 100 J/g  
 b. 100 J/g  
 c. 100 J/g  
 d. 100 J/g

8. For the reaction  $\text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$ , the  $\Delta H$  for the reaction is  $-286 \text{ kJ/mol}$  of  $\text{H}_2$  consumed. The  $\Delta H$  for the reaction is  $-286 \text{ kJ/mol}$  of  $\text{H}_2$  consumed.  
 a. 286 kJ/mol  
 b. 572 kJ/mol  
 c. 143 kJ/mol  
 d. 286 kJ/mol

9. At standard  $T_p$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  vapor has higher enthalpy than liquid water.  
 a. True  
 b. False  
 c. Not sure  
 d. None

10. The enthalpy of fusion is the lowest for:  
 a. Water at 0°C  
 b. Ice at 0°C  
 c. Liquid water at 0°C  
 d. Solid water at 0°C

11. The enthalpy of fusion is the lowest for:  
 a. Water at 0°C  
 b. Ice at 0°C  
 c. Liquid water at 0°C  
 d. Solid water at 0°C

# Multiple Choice Stoichiometry Test

**Karin Nielsen-Saines**



## Multiple Choice Stoichiometry Test:

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