

## Chemistry 2<sup>nd</sup> Semester Final Exam Review

### Chemical Bonds

1. Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.  
*A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.*
2. Use the information from the previous question to describe the shape and volume for each phase (state) of matter.  
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.  
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.  
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
3. List the three phases of matter in order of increasing intermolecular attractions.  
*A: Gas, liquid, solid.*
4. Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?  
*A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.*

### Conservation of Matter and Stoichiometry

5. The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
6. Balance the following equations:
  - a.  $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$   
*A:  $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$*
  - b.  $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
*A:  $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$*
  - c.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$   
*A:  $4\text{P} + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$*
7. What isotope is used to define one mole? *A: Carbon-12*
8. How many particles (atoms or molecules) are in one mole? *A:  $6.02 \times 10^{23}$*
9. Find the molar mass for the following chemical formulas:
  - a.  $\text{ZnF}_2$  *A: 103.377 g/mol*
  - b.  $\text{Al}_2(\text{SO}_4)_3$  *A: 342.151 g/mol*
  - c.  $\text{NH}_4\text{IO}_3$  *A: 192.941 g/mol*
10. How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
11. How many atoms are in 2075 g of He? *A:  $3.12 \times 10^{26}$  atoms*
12. What volume will 92.31 g of  $\text{CO}_2$  occupy at STP? *A: 46.985 L*
13. How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?  
 $\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$   
*A: 1.68 g NaOH*
14. In the reaction,  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ , if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?  
*A: 126 g MgO*

### Gases and Their Properties

15. What is the kinetic molecular theory (KMT)?  
*A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:*
  - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
  - These particles move in a straight line until they collide with another particle or the walls of the container.
  - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
  - There is no force of attraction between gas particles or between the particles and the walls of the container.
  - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.

# Nd Semester Final Review Questions

**Shasha Hu**



## **Nd Semester Final Review Questions:**

Digital Learning and Teaching in Chemistry Yehudit Judy Dori, Courtney Ngai, Gabriela Szeinberg, 2023-07-12 Education is always evolving and most recently has shifted to increased online or remote learning Digital Learning and Teaching in Chemistry compiles the established and emerging trends in this field specifically within the context of learning and teaching in chemistry This book shares insights about five major themes best practices for teaching and learning digitally digital learning platforms virtual visualisation and laboratory to promote learning in science digital assessment and building communities of learners and educators The authors are chemistry instructors and researchers from nine countries contributing an international perspective on digital learning and teaching in chemistry While the chapters in this book span a wide variety of topics as a whole they focus on using technology and digital platforms as a method for supporting inclusive and meaningful learning The best practices and recommendations shared by the authors are highly relevant for modern chemistry education as teaching and learning through digital methods is likely to persist Furthermore teaching chemistry digitally has the potential to bring greater equity to the field of chemistry education in terms of who has access to quality learning and this book will contribute to that goal This book will be essential reading for those working in chemical education and teaching Yehudit Judy Dori is internationally recognised formerly Dean of the Faculty of Education of Science and Technology at the Technion Israel Institute of Technology and won the 2020 NARST Distinguished Contributions to Science Education through Research Award DCRA for her exceptional research contributions Courtney Ngai and Gabriela Szeinberg are passionate researchers and practitioners in the education field Courtney Ngai is the Associate Director of the Office of Undergraduate Research and Artistry at Colorado State University Gabriela Szeinberg serves as Assistant Dean and Academic Coordinator for the College of Arts and Sciences at Washington University in St Louis Learning and Memory: A Comprehensive Reference, 2017-07-07 Learning and Memory A Comprehensive Reference Second Edition Four Volume Set is the authoritative resource for scientists and students interested in all facets of learning and memory This updated edition includes chapters that reflect the state of the art of research in this area Coverage of sleep and memory has been significantly expanded while neuromodulators in memory processing neurogenesis and epigenetics are also covered in greater detail New chapters have been included to reflect the massive increase in research into working memory and the educational relevance of memory research No other reference work covers so wide a territory and in so much depth Provides the most comprehensive and authoritative resource available on the study of learning and memory and its mechanisms Incorporates the expertise of over 150 outstanding investigators in the field providing a one stop resource of reputable information from world leading scholars with easy cross referencing of related articles to promote understanding and further research Includes further reading for each chapter that helps readers continue their research Includes a glossary of key terms that is helpful for users who are unfamiliar with neuroscience terminology Effective Classroom Turnaround John

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February 9, 1921 (*Denver Public School Monographs ; No. 5*). *The Denver Program of Curriculum Revision, 1927*, 1892  
*Enriched Teaching of Mathematics in the Junior and Senior High School* Maxie Nave Woodring, Vera Sanford, 1938  
**AIAA Aerospace Design Conference: 92-1041 - 92-1080**, 1992      *Journal of the Michigan Schoolmasters' Club*  
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## **Table of Contents Nd Semester Final Review Questions**

1. Understanding the eBook Nd Semester Final Review Questions
  - The Rise of Digital Reading Nd Semester Final Review Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Final Review Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nd Semester Final Review Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nd Semester Final Review Questions
  - Personalized Recommendations
  - Nd Semester Final Review Questions User Reviews and Ratings
  - Nd Semester Final Review Questions and Bestseller Lists
5. Accessing Nd Semester Final Review Questions Free and Paid eBooks
  - Nd Semester Final Review Questions Public Domain eBooks
  - Nd Semester Final Review Questions eBook Subscription Services
  - Nd Semester Final Review Questions Budget-Friendly Options

6. Navigating Nd Semester Final Review Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Nd Semester Final Review Questions Compatibility with Devices
  - Nd Semester Final Review Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nd Semester Final Review Questions
  - Highlighting and Note-Taking Nd Semester Final Review Questions
  - Interactive Elements Nd Semester Final Review Questions
8. Staying Engaged with Nd Semester Final Review Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nd Semester Final Review Questions
9. Balancing eBooks and Physical Books Nd Semester Final Review Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nd Semester Final Review Questions
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Nd Semester Final Review Questions
  - Setting Reading Goals Nd Semester Final Review Questions
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nd Semester Final Review Questions
  - Fact-Checking eBook Content of Nd Semester Final Review Questions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements



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