

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is requested or implied.

1. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____ in _____ and _____ in _____.
2. _____, which concentration is _____, is a _____ acid. _____, which concentration is _____, is a _____ acid. _____, which concentration is _____, is a _____ acid. _____, which concentration is _____, is a _____ acid.
3. _____, which concentration is _____, is a _____ acid. _____, which concentration is _____, is a _____ acid. _____, which concentration is _____, is a _____ acid.

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19 2 Hydrogen Ions And Acidity Worksheet Answers

N Noddings



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