

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. neutral solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the values indicated that are indicated in the space provided. Be sure to include units when appropriate.

1. What substance is _____, contains hydrogen ions?
2. _____ is the concentration of those acids present in rainwater that equal to _____.
3. _____ is the concentration of those acids present in _____.
4. _____ is the concentration of those acids present in _____.
5. _____ is the concentration of those acids present in _____.
6. _____ is the concentration of those acids present in _____.
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17. _____ is the concentration of those acids present in _____.
18. _____ is the concentration of those acids present in _____.
19. _____ is the concentration of those acids present in _____.
20. _____ is the concentration of those acids present in _____.

1. _____
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17. _____
18. _____
19. _____
20. _____

19 2 Hydrogen Ions And Acidity Answer Key

Terry C. Jones



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