

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate formula for each of the following acids, bases, and salts, showing dissociation reactions, and calculate the pH of each solution.

1. Acetic acid,  $CH_3COOH$ ,  $K_a = 1.8 \times 10^{-5}$
2. Hydrofluoric acid,  $HF$ ,  $K_a = 6.6 \times 10^{-4}$
3. Hydrocyanic acid,  $HCN$ ,  $K_a = 4.9 \times 10^{-10}$
4. Hydrobromic acid,  $HBr$ ,  $K_a = 1.0 \times 10^3$
5. Hydrochloric acid,  $HCl$ ,  $K_a = 1.0 \times 10^7$
6. Hydroiodic acid,  $HI$ ,  $K_a = 1.0 \times 10^9$
7. Hydroperchloric acid,  $HClO_4$ ,  $K_a = 1.0 \times 10^4$
8. Hydroperbromic acid,  $HBrO_4$ ,  $K_a = 1.0 \times 10^4$
9. Hydroperchloric acid,  $HClO_4$ ,  $K_a = 1.0 \times 10^4$
10. Hydroperbromic acid,  $HBrO_4$ ,  $K_a = 1.0 \times 10^4$
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1. Acetic acid,  $CH_3COOH$
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19. Hydroperchloric acid,  $HClO_4$
20. Hydroperbromic acid,  $HBrO_4$

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**KJ Lindholm-Leary**



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