

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                       |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution      |
| 4. autoionization | 5. acid rain                | 6. acid rain solution |
|                   | 7. acid solution            | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water.
3. Write the equation for the autoionization of water.
4. Write the equation for the autoionization of water.
5. Write the equation for the autoionization of water.
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17. Write the equation for the autoionization of water.
18. Write the equation for the autoionization of water.
19. Write the equation for the autoionization of water.
20. Write the equation for the autoionization of water.

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
5.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
6.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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19.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
20.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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