

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water which produces an acidic solution.
3. Write the equation for the autoionization of water which produces a basic solution.
4. Write the equation for the autoionization of water which produces a neutral solution.
5. Write the equation for the autoionization of water which produces an acidic solution.
6. Write the equation for the autoionization of water which produces a basic solution.
7. Write the equation for the autoionization of water which produces a neutral solution.
8. Write the equation for the autoionization of water which produces an acidic solution.
9. Write the equation for the autoionization of water which produces a basic solution.
10. Write the equation for the autoionization of water which produces a neutral solution.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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10. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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