

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                    |                                  |
|------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the acids and bases that are indicated in the partial equations below. Complete each acid or base ionization equation.

1. When autoionized, \_\_\_\_\_, releases hydrogen ions.
2. Autoionization of \_\_\_\_\_, the autoionization of \_\_\_\_\_, produces an acidic solution.
3. \_\_\_\_\_, when autoionized, forms \_\_\_\_\_, an acid.
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