

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- k. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- l. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H<sub>2</sub>SO<sub>4</sub>**.

4. **HClO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{ClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{ClO}_4 + \text{H}_3\text{O}^+$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{H}_3\text{ClO}_4$  in a acidic acid.

c. Dissolve **1.00\text{M}**  $\text{NaClO}_4$  in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}**  $\text{H}_3\text{O}^+$  in a portion of water.

d. In your **reaction** prediction, is  $\text{HClO}_4$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because  $\text{HClO}_4$  is a **strong** acid.

e. Identify the conjugate base of  $\text{HClO}_4$ :  **$\text{ClO}_4^-$**

f. Identify  $\text{NaClO}_4$  as **strong**, **weak**, or **none**: **strong**. Because it contains a conjugate base  $\text{ClO}_4^-$ .

5.  $\text{H}^+$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_3\text{O}^+ + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{O}^+ + \text{H}_2\text{O}$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{H}_3\text{O}^+$  in a acidic base.

c. Dissolve **1.00\text{M}**  $\text{H}_3\text{O}^+$  in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}**  $\text{H}_2\text{O}^+$  in a portion of water.

d. In your **reaction** prediction, is  $\text{H}_3\text{O}^+$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because  $\text{H}_3\text{O}^+$  is a **strong** acid.

e. Identify the conjugate base of  $\text{H}_3\text{O}^+$ :  **$\text{H}_2\text{O}$**

f. Identify  $\text{H}_3\text{O}^+$  as **strong**, **weak**, or **none**: **strong**

6. **H<sub>2</sub>SO<sub>4</sub>** will react with water and lose a  $\text{H}^+$  ion in the process.

a. Write a balanced equation for this reaction:  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{SO}_4^+ + \text{H}_3\text{O}^+$

b. Write a  $\text{pH}$  response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating  $\text{H}_2\text{SO}_4$  in a acidic acid.

c. Dissolve **1.00\text{M}**  $\text{NaHSO}_4$  in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}**  $\text{H}_3\text{O}^+$  in a portion of water.

d. In your **reaction** prediction, is  $\text{H}_2\text{SO}_4$  acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because  $\text{H}_2\text{SO}_4$  is a **strong** acid.

e. Identify the conjugate base of  $\text{H}_2\text{SO}_4$ :  **$\text{HSO}_4^-$**

f. Identify  $\text{NaHSO}_4$  as **strong**, **weak**, or **none**: **strong**. Because it contains a conjugate base  $\text{HSO}_4^-$  in the **strong** acid  $\text{H}_2\text{SO}_4$ .

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**Henry Watts**

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