

**2****MATTER AND CHANGE****SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)**

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

**► Describing Matter (page 39)**

1. The mass of an object is a measure of the amount of matter the object contains.
2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

**► Identifying Substances (page 40)**

3. Matter that has a uniform and definite composition is called a substance.
4. Is the following sentence true or false? All samples of a substance have different physical properties. false
5. A physical property is a quality or condition of a substance that can be observed or measured without changing the substance's composition.
6. Circle the letter of the term that is NOT a physical property.
  - a. hardness
  - b. color
  - c. boiling point
  - d. melting
7. Look at Table 2.1 on page 40. What is the melting point of bromine? -7°C
8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.
  - a. sulfur
  - b. chlorine
  - c. gold
  - d. copper

## Section 3 1 Chemistry Answers

**Henry Watts**

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*Micro-chemistry of Poisons* Theodore George Wormley, 1885      **Russian Journal of General Chemistry**, 1998      **An**

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**Pharmaceutical Journal**, 1896      **A Dictionary of Chemistry and the Allied Branches of Other Sciences** Henry Watts, 1879

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## Table of Contents Section 3 1 Chemistry Answers

1. Understanding the eBook Section 3 1 Chemistry Answers
  - The Rise of Digital Reading Section 3 1 Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 3 1 Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 3 1 Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 3 1 Chemistry Answers
  - Personalized Recommendations
  - Section 3 1 Chemistry Answers User Reviews and Ratings
  - Section 3 1 Chemistry Answers and Bestseller Lists
5. Accessing Section 3 1 Chemistry Answers Free and Paid eBooks
  - Section 3 1 Chemistry Answers Public Domain eBooks
  - Section 3 1 Chemistry Answers eBook Subscription Services
  - Section 3 1 Chemistry Answers Budget-Friendly Options
6. Navigating Section 3 1 Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 3 1 Chemistry Answers Compatibility with Devices
  - Section 3 1 Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 3 1 Chemistry Answers
  - Highlighting and Note-Taking Section 3 1 Chemistry Answers
  - Interactive Elements Section 3 1 Chemistry Answers
8. Staying Engaged with Section 3 1 Chemistry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Section 3 1 Chemistry Answers

9. Balancing eBooks and Physical Books Section 3 1 Chemistry Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Section 3 1 Chemistry Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Section 3 1 Chemistry Answers

- Setting Reading Goals Section 3 1 Chemistry Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Section 3 1 Chemistry Answers

- Fact-Checking eBook Content of Section 3 1 Chemistry Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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