



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Calculations

Use the following equations to calculate the pH of a solution with a hydrogen ion concentration of $1.0 \times 10^{-4} \text{ M}$.

$\text{pH} = -\log_{10}[\text{H}^+]$ $\text{pH} = -\log_{10}[\text{H}^+]$
concentration concentration

Concentration	pH
$1.0 \times 10^{-4} \text{ M}$	4.00
$1.0 \times 10^{-5} \text{ M}$	5.00
$1.0 \times 10^{-6} \text{ M}$	6.00
$1.0 \times 10^{-7} \text{ M}$	7.00
$1.0 \times 10^{-8} \text{ M}$	8.00
$1.0 \times 10^{-9} \text{ M}$	9.00
$1.0 \times 10^{-10} \text{ M}$	10.00
$1.0 \times 10^{-11} \text{ M}$	11.00
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