

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a solution and the pH of a solution based on concentration.
2. Calculate pH and concentration of acids and bases and the pH of a solution based on concentration.
3. Calculate the pH of a solution (25°C).

Vocabulary

- | | | |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. acid rain |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the pH of a solution that is 0.001 M H^+ ?

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