

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                            |                   |                   |
|----------------------------|-------------------|-------------------|
| acid dissociation constant | 1. conjugate base | 2. weak acid      |
| acid ionization            | 3. conjugate acid | 3. autoionization |
| acid ionization constant   | 4. weak base      |                   |

## Key Equations

- |                                    |                                   |
|------------------------------------|-----------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[HB][OH^-]}{[B]}$ |
| 3. $K_a + K_b = K_w$               |                                   |

## Part A Completion

Write the appropriate words in the blanks to complete the following statements. Write the chemical formula of the acid or base in the blank.

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
2. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
3. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
4. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
5. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
6. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
7. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
8. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .

9. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
10. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
11. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
12. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
13. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
14. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
15. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
16. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .

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**Valerie C. Scanlon, Tina Sanders**



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