

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

- Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
- Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
Gas particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
- List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
- Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

- The atomic mass unit and the mole are defined using what isotope?
A: Carbon-12
- Balance the following equations:
 - $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$
A: $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$
 - $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
A: $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 8\text{ CO}_2 + 6\text{H}_2\text{O}$
- What isotope is used to define one mole?
A: Carbon-12
- How many particles (atoms or molecules) are in one mole?
A: 6.02×10^{23}
- Find the molar mass for the following chemical formulas:
 - ZnF_2
A: 103.377 g/mol
 - $\text{Al}_2(\text{SO}_4)_3$
A: 342.151 g/mol
- How many moles are in 4.37 kg of NaOH?
A: 109.3 mol
- How many atoms are in 2075 g of He?
A: 3.12×10^{26} atoms
- What volume will 92.31 g of CO₂ occupy at STP?
A: 46.985 L
- How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
$$\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba}(\text{OH})_2$$

A: 1.68 g NaOH
- In the reaction, $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

- What is the kinetic molecular theory (KMT)?
*A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gas particles are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
 - These particles move in a straight line until they collide with another particle or the walls of the container.
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
 - There is no force of attraction between gas particles or between the particles and the walls of the container.
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.*

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