

TEST YOUR SELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1_1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.

b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.

c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-4} g or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebound of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

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Forensic Science Jay A. Siegel, Kathy Mirakovits, 2006-09-07 Forensic Science The Basics explains every aspects of crime scene investigation moving from basic areas of criminalistics and beyond to pathology anthropology and engineering It also explores new and emerging areas such as forensic entomology With no previous knowledge of either science or law required information is self contained and conveyed at the lowest possible non scientific level making this text suitable for both lower level academic adoptions as well as for a general audience It also offers a complete package of ancillary material for instructors Comprehensive and Up to Date Covers DNA drugs firearms fingerprints and trace evidence Includes cutting edge material on spectroscopy chromatography microscopy odontology and entomology Demonstrates the practical application of modern chemistry biology and other laboratory sciences Each chapter Opens with learning objectives a chapter outline and an introduction Closes with a summary and review questions for self testing Contains real life examples many from the author s own experience Build an exceptional classroom experience with this dynamic resource More than 200 full color nongraphic illustrations Countless figures tables and charts A wealth of supporting material including lecture slides and test questions available on www.classwire.com Real case studies to demonstrate forensic concepts in action Suggested student projects to reinforce learning Appropriate for High School and University Students Written in the lucid and concise style of a master teacher Fully explains the scientific basics required Omits potentially traumatic photographs and subject matter About the Author Eminently qualified to create this work Jay Siegel is both a practicing forensic expert and a master instructor He has worked for the Virginia Bureau of Forensic Sciences and published extensively in the field He continues to be called upon as an expert witness having testified over 200 times in state federal and military courts across the country With nearly thirty years of teaching experience he is highly active in curriculum development for forensic science classes taught at all levels from junior high through graduate school He is currently director of the Forensic and Investigative Sciences Program at Purdue University in Indiana In February of 2009 Mr Siegel received the Distinguished Fellow award from the American Academy of Forensic Sciences at its annual meeting This is the highest honor that the Academy bestows upon a fellow In addition George Washington University has selected Mr Siegel for the 2008 2009 Distinguished Alumni Scholar This award the highest that the University bestows upon its alumni is designated for those who have made truly outstanding contributions to the knowledge base of their disciplines For Instructors Only Develop and Customize Your Curriculum Draw from hundreds of PowerPoint slides and illustrations to supplement your lectures Organize your class with Dr Siegel s helpful outlines and learning objectives Review answers to end of chapter questions Build exams for different levels from a giant test bank of problems This book also works in conjunction with Forensic Science Laboratory Manual and Workbook Revised Edition All ancillary material will be available in convenient website format at www.classwire.com Upon request photographs lecture slides and a test bank are also available to instructors on CD **Fachenglisch für**

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Quiz Refresher in Chemistry Rolland J. Gladieux, 1953

Study Guide for Chemical Principles, Fourth Edition, by

Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984 Timetable University of Illinois at Urbana-Champaign, 2000 Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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English Mechanic and World of Science, 1879 **Report[s], [minutes of Evidence, Indexes, Answers to Questions]**. Great Britain. Royal Commission on Labour, 1894 **The Journal of Education**, 1899 A Complete Preparation for the MCAT Aftab S. Hassan, James L. Flowers, 1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress

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