

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization          | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $pH$                     |                   |

## Key Equation

- |                       |                       |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes.  $K_w$  is the ionization constant of water.  $K_w$  is the concentration of  $H_3O^+$  and  $OH^-$  in pure water at  $25^\circ C$ .
2. Hydrochloric acid dissociates from  $HCl$  to  $H^+$  and  $Cl^-$ .  $K_a$  is the ionization constant of  $HCl$ .  $K_a$  is the concentration of  $H^+$  and  $Cl^-$  in pure water at  $25^\circ C$ .
3.  $K_b$  is the ionization constant of  $OH^-$ .  $K_b$  is the concentration of  $OH^-$  and  $H^+$  in pure water at  $25^\circ C$ .
4.  $K_a$  is the ionization constant of  $HCl$ .  $K_a$  is the concentration of  $H^+$  and  $Cl^-$  in pure water at  $25^\circ C$ .

5.  $K_b$  is the ionization constant of  $OH^-$ .
6.  $K_a$  is the ionization constant of  $HCl$ .
7.  $K_b$  is the ionization constant of  $OH^-$ .
8.  $K_a$  is the ionization constant of  $HCl$ .
9.  $K_b$  is the ionization constant of  $OH^-$ .
10.  $K_a$  is the ionization constant of  $HCl$ .
11.  $K_b$  is the ionization constant of  $OH^-$ .
12.  $K_a$  is the ionization constant of  $HCl$ .
13.  $K_b$  is the ionization constant of  $OH^-$ .
14.  $K_a$  is the ionization constant of  $HCl$ .
15.  $K_b$  is the ionization constant of  $OH^-$ .
16.  $K_a$  is the ionization constant of  $HCl$ .

# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Advisory Council on College Chemistry**



## 19 2 Hydrogen Ions And Acidity Worksheet Answers:

*Class 10th Science Worksheet*, This book is as per the guidelines syllabus and marking scheme issued by CBSE for Class X The salient features of this workbook are The questions in the this book have been so designed that complete syllabus is covered This book help students to identify their weak areas and improve them Additional it will help students gain confidence The questions in the book are of varying difficulty level and will help students evaluate their reasoning analysis and understanding of the subject matter **Topics-aids** Rod O'Connor,1975 **Maple® for Environmental Sciences** Bill Scott,2012-12-06 What is this book about Please take this book as it is a working document It started as an idea that has grown It will never be correct but should be self correcting In the limit if there is one the book should approach a correct state It is not the detail and the numbers that matter but the structures and the order These structures are inherently linked with the many minds that have made Maple the minds of perhaps the best mathematicians certainly some of the most useful Our environment is not separate from mathematics mathematics is but one tool of several to help with understanding the environment It is a harsh tool that requires numbers and symbolism Maple handles the symbolism superbly numbers need more consideration We have included a substantial amount on reading and writing numbers data and dealing with floating point numbers It is the devil in the detail that continually comes back to us in working with Mathematics and Maple It becomes raw and defined Many of the things we do have rational and logical bases but we don't know what they are Often in following the code and talking with an input line to Maple the detailed way of performing a task becomes clear But not without frustration the task is invariably simple though [Educart ICSE Class 10 One-shot Question Bank 2026 Chemistry \(strictly for 2025-26 boards\)](#) Sir Tarun Rupani,2025-07-12 Fast track your Chemistry revision with this exam ready resource This One shot Question Bank by Sir Tarun Rupani is designed to help ICSE Class 10 students revise the complete Chemistry syllabus quickly and thoroughly It simplifies theory boosts numerical accuracy and ensures strong exam practice all aligned with the 2025 26 ICSE syllabus Key Features Strictly Based on ICSE 2025 26 Curriculum Complete chapter coverage including Periodic Table Chemical Bonding Acid Base Organic Chemistry and more One shot Format Each chapter includes concise concept notes chemical equations reactions and key diagrams for quick recall Complete Coverage of Question Types Includes objective short long answers equation based numerical and reasoning questions Chapterwise PYQs Included Practice with previous years ICSE board questions to understand trends and improve retention Solved Answers in ICSE Format Clear well structured solutions using proper units chemical symbols and balanced equations Smart Revision Focus Special tips to avoid common mistakes in writing reactions balancing equations and attempting numericals Why Choose This Book This Chemistry One shot by Sir Tarun Rupani is built for smart preparation whether you're revising at the last minute or practising throughout the term It helps you approach each question with clarity confidence and the precision needed to score high in the 2026 ICSE board exam *Educart One-shot Science CBSE Class 10 Question Bank 2025-26 on new*

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